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Problems Key Section 3.1 Solids, Liquids, and Gases For each of the questions in this section, illustrate your written answers with simple drawings of the particles that form the structures of the substances mentioned. You do not need to be specific about the nature of the particles. Think of them as simple spheres, and draw them as circles. 33.

Chapter 3 The Structure of Matter and the Chemical Elements

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Solutions Manual Chemistry: Matter and Change • Chapter 12 239 CHAPTER 12 SOLUTIONS MANUAL Section 12.3 Liquids and Solids pages 415-424 Section 12.3 Assessment page 424 18. Contrast the arrangement of particles in solids and liquids. The particles are closer together in solids than in liquids because of intermolecular attractions.

States of Matter

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to 38.9C? What is the energy value of the olive oil in kJ/g and kcal/g?

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Solutions Manual Chemistry: Matter and Change • Chapter 8 121 Section 8.1 The Covalent Bond pages 240–247 Practice Problems page 244 Draw the Lewis structure for each molecule. 1. PH 3 H HH H— H H P respectively, for single, double, and triple P — - 2. H 2 S H H H — H S S — 3. HCl H Cl H — Cl 4. CCl 4 Cl Cl Cl C C — — Cl Cl ...

Covalent BondingCovalent Bonding

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Section 2.3 Density and Density Calculations Goal: To describe what density is, how it can be used as a conversion factor, and how density can be calculated. Density calculations are common in chemistry. The examples in this section show you how these calculations are done, and perhaps more important, these density calculations provide

Chapter 2 Unit Conversions
Page 8/11

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5.2: Wavelength and Frequency Calculations - Chemistry ...

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